



The Effect of Nanoparticles with Sorption Capacity on the Bioaccumulation of Divalent Ions by Aquatic Plants

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Abstract

Wide utilization of nanotechnologies results in the release of nanomaterials into the environment. Considering the sorption capacity of nanoparticles (NPs) and their growing concentration in aquatic environment, the modification of metal ion uptake by direct NP–metal ion reaction, including adsorption of ions onto the surface of the NPs, can be assumed. The aim of the study was to characterize the sorption capacity of zinc and titanium oxide nanoparticles (TiO₂ NPs 25 and 100 nm; ZnO NPs 50 and 100 nm) and to examine the bioaccumulation of Zn(II) and Sr(II) ions by aquatic plants *Salvinia natans* and *Elodea canadensis* in the presence of the above-mentioned NPs. The results have shown that NPs can adsorb metal ions present in the water. It was also observed that the sorption efficiency strongly depends on the: (1) concentrations of NPs and increase in the range 0.01–5 g L⁻¹; (2) pH—the highest sorption efficiency was observed at pH 8; (3) chemical composition of the medium (presence of humic substances at 0.01 g L⁻¹ increased the sorption efficiency of Sr on TiO₂ 25 nm from 10 to 50%). Under some conducive conditions, even 90% of Zn(II) ions can be removed from a solution through sorption using ZnO nanoparticles. Bioaccumulation of Zn(II) and Sr(II) ions by *S. natans* and *E. canadensis* was modified in the presence of investigated NPs. Bioaccumulation of Sr in *E. canadensis* and *S. natans* was reduced by NPs of ZnO 100 nm at concentration 1 g L⁻¹ of about 49 and 23%, respectively. Bioavailability of Zn was the most affected by NPs of ZnO 100 nm at 1 g L⁻¹. Bioaccumulation of Zn in *E. canadensis* was reduced by 68%. An opposite effect was found for *S. natans*—addition of ZnO NPs caused an increase of Zn bioaccumulation by 45%.

Keywords TiO₂ nanoparticles · ZnO nanoparticles · Sorption · Aquatic plants · *Salvinia natans*, *Elodea canadensis*

Introduction

Achievements in the area of nanoscience and nanotechnology find practical applications in many fields of the industry and in daily life. Nanoparticles (NPs), up to 100 nm large, compared to other materials, have unique physicochemical properties resulting from a high ratio of surface area to volume or weight and exhibit unusual interactions with the media (Alkilany and Murphy 2010; Bystrzejewska-Piotrowska et al. 2009). One of the side effects of wide utilization of nanotechnologies is the release of nanomaterials into the environment. Therefore,

there is a growing need to study the diverse aspects of the presence of NPs in the environment. The focus of the researchers was mainly on the toxicity of NPs toward various kinds of organisms from a biomedical (Zoroddu et al. 2014; Yang et al. 2016) and ecotoxicological (Baker et al. 2014; Ma et al. 2010; Książyk et al. 2015; Thwala et al. 2016) perspective. Some studies are focused on the bioaccumulation of NPs by organisms (Bystrzejewska-Piotrowska et al. 2012a; Asztemborska et al. 2014; Asztemborska et al. 2015a, b; Jakubiak et al. 2014; Jiménez-Lamana et al. 2016). However, investigation of direct interaction of NPs with living organisms does not provide complete information on real threat to the environment. Besides the toxicity, the influence of NPs on the chemical and physical behavior of soil or water components must also be considered. One of the possible effects caused by NPs is modification in the uptake of metal ions and other

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substances. It can be achieved by direct NP–metal ion reaction, including adsorption of ions onto the surface of NPs. Another effect is the influence on the transport systems of organisms by forming new, large-sized pores and routes, which increases the uptake of compounds or their adsorption onto the surface of the roots, resulting in inhibition of the uptake. The effect of NPs on the uptake of xenobiotics was documented in the literature. Baun et al. (2009) have found that the aggregates of nano-C60 (Buckminsterfullerenes) influence the toxicity of organic compounds for algae *Pseudokirchneriella subcapitata* and crustaceans *Daphnia magna*. Bioaccumulation and toxicity of cadmium, zinc (Tan and Wang 2014), and copper (Fan et al. 2011) by *D. magna* are modified by titanium dioxide NPs. Nanoparticles of TiO₂ enhance the uptake of phenanthrene by *Scapharca subcrenata* (Tian et al. 2014) and the accumulation of cadmium (Zhang et al. 2007) and arsenate (Sun et al. 2007) in carp.

One of the possible explanations for the modification of bioaccumulation by NPs is their sorption capacity. For example, carbon nanotubes show a high capacity for the adsorption of toxic metals such as Pb²⁺, Cu²⁺, and Cd²⁺ (Li et al. 2003). Saleh (2016) have demonstrated the ability of nanocomposites of carbon nanotubes and silica nanoparticles for the adsorption of Pb(II). Magnetite nanoparticles were investigated by Giraldo et al. 2013 as sorbents for Pb(II), Cu(II), Zn(II), and Mn(II). Adsorption was strongly influenced by pH and temperature, and the effect depended on the metal ion considered. Modified magnetite nanoparticles were applied for the adsorption of precious metals Pd(IV), Au(III), Pd(II), Ag(I) (Yen et al. 2017), Cu(II) (Liu et al. 2016), and Hg(II) (Wang et al. 2016a, b). Mesoporous NiO nanoparticles find application in the adsorption of Cr(VI) (Behnajady and Bimeghdar, 2014). Mahdavi et al. (2013) investigated on the removal of heavy metal ions (Cd(II), Cu(II), Ni(II), and Pb(II)) by oxide metal nanoparticles: TiO₂, MgO, and Al₂O₃. Among them, MgO nanoparticles were indicated as the most promising sorbents. Adsorption of divalent ions of Pb, Cd, Cu, Ni, and Zn to TiO₂ nanoparticles and TiO₂ bulk particles was studied as a function of the particle size, concentration, and exhaustion by Engates and Shipley (2011). Results confirmed that TiO₂ nanoparticles are good sorbents of metal ions; during the first-adsorption cycle, 100% of Pb, 99% of Cd and Zn, and 68% Cu were adsorbed. Utilization of ZnO nanoparticles as an adsorbent of Zn(II), Cd(II), and Hg(II) was examined by Sheela et al. (2012). The sorption capacities of investigated NPs were 357, 387, and 714 for Zn, Cd, and Hg, respectively. Asztemborska et al. (2016) have shown that alumina nanoparticles (Al₂O₃ NPs) have the ability to adsorb Sr, and to a lesser extent, cesium (sorption efficacy was 40 and 20% for Sr and Cs, respectively). However, contrary to expectations, Al₂O₃

NPs did not affect the mycoextraction of radioisotopes by the mycelia of *Pleurotus eryngii*. In a medium containing *P. eryngii* mycelia and NPs, concurrent processes, such as accumulation of radionuclides, accumulation of NPs, and sorption of radionuclides on the nanoparticle surfaces, can be distinguished.

Identification of aquatic environment as one of the most exposed to contamination with NPs enforces the extensive study of various effects of nano-level substances with organisms. One of the possible interactions is the modification of ion uptake by organisms as an effect of the sorption capacity of NPs. The aims of the study were: (1) to characterize the sorption capacities of zinc and titanium oxide nanoparticles (to find the influence of particle size and concentration, sorption temperature, pH, and time on the adsorption efficiency) and (2) to examine the bioaccumulation of Zn(II) and Sr(II) ions by *S. natans* and *Elodea canadensis* in the presence of metal oxide nanoparticles.

Materials and Methods

Reagents

Nanoparticles of TiO₂ (particle sizes 25 nm (surface area > 14.0 m²/g) and 100 nm (surface area 45–55 m²/g)) and ZnO [particle sizes 50 nm (surface area > 10.8 m²/g) and 100 nm (surface area 15–25 m²/g)] were purchased from Sigma-Aldrich. In more detail, nanoparticles were characterized elsewhere (Baek and An 2011; Sadiq et al. 2011; Bystrzejewska-Piotrowska et al. 2012b). The stock solution was prepared by mixing 100 mg of solid NPs and 100 mL of deionized water to obtain a final concentration of 1 g L⁻¹. To achieve the concentration required for the experiments, a proper dilution of the stock solution was done. All solutions were freshly prepared and sonicated before experiments.

The Zn and Sr stock solutions were obtained by dissolving 100 mg of ZnCl₂ or SrCl₂ in water (100 mL) to obtain a final concentration of 1 g L⁻¹. Radioactive isotopes of Zn (⁶⁵Zn) and Sr (⁹⁰Sr) with activities 0.1 and 1 MBq, respectively, were purchased from Polatom (Otwock, Poland) and these were used as isotope tracers. Zn and Sr were present in all experimental variants at a constant concentration of 0.01 mg L⁻¹ with an activity of 1 kBq L⁻¹.

Adsorption Experiments

To characterize the sorption of Zn and Sr by NPs, experiments were performed under controlled conditions. In a 50-mL vessel, a fixed amount of metal ions and an

appropriate amount of NPs were added; the final volume of the experimental solution was 10 mL. All experimental solutions were continuously agitated on an orbital shaker (120 rpm, 25 °C). During the first step of experiment, several different concentrations of NPs were applied: 0.01–8 g L⁻¹. For selected NP concentration (0.1 and 1 g L⁻¹), sorption experiments were performed at variable time (0–24 h). Next, using 1 g L⁻¹ NPs and sorption time of 15 min, the influence of temperature (5–90 °C) and pH (2–10) on sorption efficiency was investigated. Finally, sorption experiments were performed with the addition of 0.01 g L⁻¹ humic acids (Sigma-Aldrich; solution in deionized water) to evaluate the influence of organic matter on the sorption capacity of NPs (sorption time 15 min and 24 h). All experimental series included reference samples, containing Zn or Sr ions without any NPs, used for the calculation of sorption capacity of NPs. After finishing the sorption, solutions were centrifuged (30 min; 16,000 rpm) and analyzed for the radioactivity of radioisotopes.

Bioaccumulation Studies

Plants used for bioaccumulation studies—adult *S. natans* and *E. canadensis*—obtained from the Botanical Garden of the Warsaw University, were cultivated in containers with solution containing metal oxide nanoparticles at 0.1 and 1 g L⁻¹ and Zn or Sr at 0.01 mg L⁻¹ under controlled conditions for 8 days. Reference cultivation with Zn or Sr ions without any NPs was carried out in parallel. *Salvinia natans* plants were cultivated in containers with 20 mL of experimental solution (five plants per container). The cultivation of *E. canadensis* was conducted in 100 mL of solution (one plant per container). After finishing the bioaccumulation experiments, plants were rinsed with deionized water, dried, homogenized, and analyzed for the radioactivity of radioisotopes.

Radioactivity Measurements

The activities of Zn in powdered plant samples of mycelium were determined using a low-background coaxial high-purity germanium detector (model GC1520; Canberra Industries, Zellik, Belgium). The relative efficiency of this detector for the ⁶⁰Co background lines at 1332 keV was 18.0%, the resolution (FWHM) was 1.74 keV, and the measurement time of each sample continued for over 3 h. Energy and efficiency calibrations were carried out using a reference material from the International Atomic Energy Agency (Vienna, Austria): IAEA-156. The activities of Sr were measured in a Geiger–Muller detector equipped with an SPF-32 plastic scintillation probe, and 5-cm-thick lead housing. Measurements of each sample were carried out within 10 min and these were referred to reference material

with a known activity of ⁹⁰Sr. The activities of liquid samples were measured in a Beckman LS6000 scintillation counter. For this purpose, 1 mL of sample was diluted with 5 mL of scintillation cocktail Rotiszint eco plus (Carl Roth GmbH), placed in a 20-mL plastic vial (Sarstedt) and counted within 10 min in CPM (counts per minute).

Data Analysis

All experiments were performed in at least three replicates. Based on the activities of standards and reference samples, sorption efficiency was calculated. Analysis of variance (ANOVA) and Pearson correlation coefficient (significance level < 0.05) were used to estimate differences in the sorption efficiency or bioaccumulation of investigated ions between experimental variants. Values were considered as different when $p < 0.05$.

Results and Discussion

Characteristics of Sorption: Effect of NPs Concentration

To characterize the sorption capacity of NPs, the sorption efficiency depending on the size and concentration of NPs, time, temperature, pH, and presence of organic matter was investigated as the first step of the study. The dependence of Zn and Sr adsorption efficiencies of different-sized ZnO (50 and 100 nm) and TiO₂ (25 and 100 nm) NPs on the concentration of NPs is shown in Fig. 1. Sorption efficiencies of Zn and Sr differ for investigated nanoparticles. Between the two ions, the lowest sorption efficiency (max. 13%) was found for 25 nm TiO₂ NPs. Large-sized NPs of TiO₂ (100 nm) have shown better sorption capacity toward investigated ions and reach the value of 48 and 37% for Sr and Zn, respectively. Some differences were found in the sorption of investigated ions in the case of ZnO NPs. For both Sr and Zn, the sorption efficiency was size-independent, in practice. Nanoparticles of ZnO remove about 40–45% of strontium and nearly 90% of zinc from the solution. Some different mechanisms of metal ion sorption as ion exchange or adsorption may occur. One of the possibilities is that the ions replace exchangeable cations, for example, hydroxyl groups. Such mechanism was postulated for sorption of Zn(II), Cd(II), and Hg(II) onto ZnO nanoparticles (Sheela et al. 2012). Difference in physical and chemical properties of Sr and Zn will explain the observed differences in sorption efficiency.

In accordance with the expectation in the range of low-studied concentrations (0.01–1 g L⁻¹), an increase of Sr and Zn sorption efficiency with an increase in the concentration of TiO₂ 100 nm, ZnO 50 nm, and ZnO 100 nm

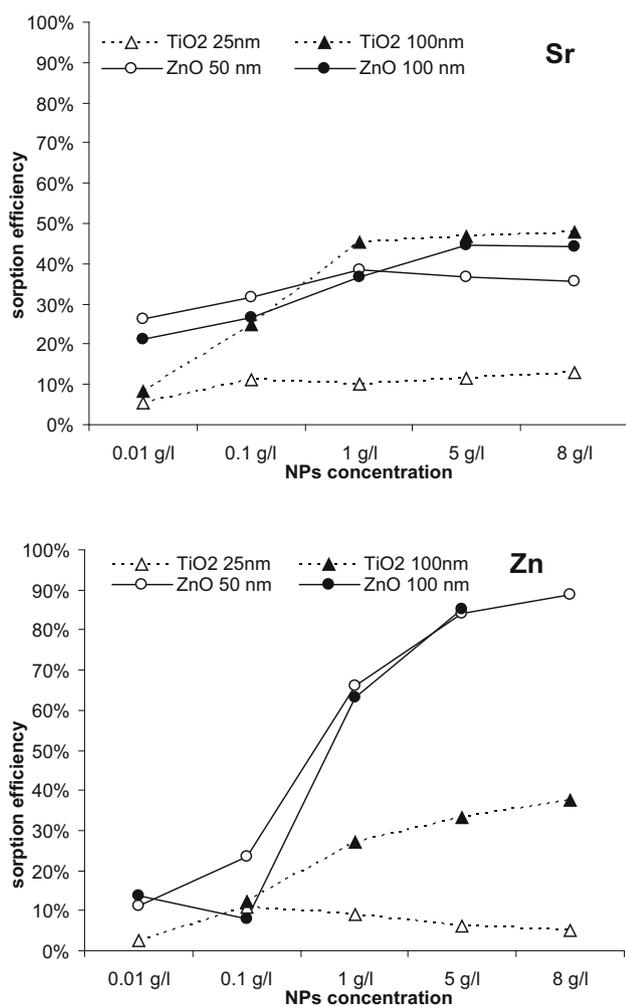


Fig. 1 Sorption efficiency of Sr and Zn for different concentrations of TiO₂ 25 nm, TiO₂ 100 nm, ZnO 50 nm, and ZnO 100 nm NPs (SD ≤ 5%)

NPs is observed; with increasing sorbent concentration, the available surface sites increases. Similar relationship was found by Engates and Shipley (2011). However, at higher concentrations of TiO₂ 100 nm, ZnO 50 nm, and ZnO 100 nm NPs (> 1 g L⁻¹), the sorption of Sr was independent of the sorbent concentration. Sorption is controlled by the nature of the adsorbent and charge of ions and availability. We can expect that at some relatively higher concentrations, the surface properties of NPs are changed, preventing more effective sorption of ions. Additionally, aggregation in concentrated solutions is much more probable and it results in the reduction of surface available for metal ions. In solution of NPs with relatively higher density, some limitations in the diffusion of ions are also possible. More physicochemical studies at particle level must be undertaken in the future to find the explanation for the observed phenomenon. Different observation was made for Zn sorption at a higher concentration of TiO₂ 100 nm,

ZnO 50, and ZnO 100 nm NPs. The lowest sorption efficiency was observed for TiO₂ 100 nm NPs in comparison with both ZnO NPs; however, the sorption efficiency for the mentioned NPs increases with their concentration. For the smallest investigated NPs, TiO₂ 25 nm, the sorption capacity was least efficient and some dependence on the concentration was observed.

Characteristics of Sorption: Effect of Sorption Time and Temperature

In contrast to concentration and particle size, sorption time in the studied range (0–24 h) was not a factor affecting the sorption efficiency. The percentage of adsorbed ions was the same after 24 h of experiments when the NPs were separated from ions just after mixing (sorption time < 60 s was limited only by the time required for laboratory operations). Thus, adsorption is the only process, which occurred rapidly. Similar observation was made by Engates and Shipley (2011), while investigating sorption of some divalent metal ions on ZnO NPs.

The adsorption of metal ions was also examined with regard to the function of temperature (Fig. 2), which is a very important factor affecting the sorption processes. The experimental results showed that Sr sorption onto TiO₂ 100 nm, ZnO 50, and ZnO 100 nm NPs does not depend on the temperature in the studied range 5–90 °C. Only for the smallest NPs of interest—TiO₂ 25 nm—the percentage of adsorbed Sr ions increased twice, when the temperature changed from 30 to 90 °C. In contrast to Sr, the sorption of Zn ions increased with the temperature for all investigated NPs, indicating that the adsorption of metal ions onto ZnO is an endothermic process. This might attribute to the increased ion mobility and collision efficiency between the ions and the nanoparticles with the increasing temperature. De-protonation reactions were easy at high temperature and the amount of non-protonated groups on the nanoparticle surface increased. The results are contrary to the data published by Sheela et al. (2012); however, the size and preparation of the NPs and the sorption conditions were different, which can result in different observations.

Characteristics of Sorption: Effect of pH

One of the most important factors affecting the chemistry of both, NPs as an adsorbent and metal ion, is the pH of the solution. Sorption efficiency of Zn and Sr was determined under acidic, neutral, and alkaline conditions (Fig. 3). It was assumed that the sorption capacity of NPs increased with pH within the range 2–8. At the highest investigated pH 10, some differences between sorption of Sr and Zn were observed. The percentage of Zn and Sr ions adsorbed

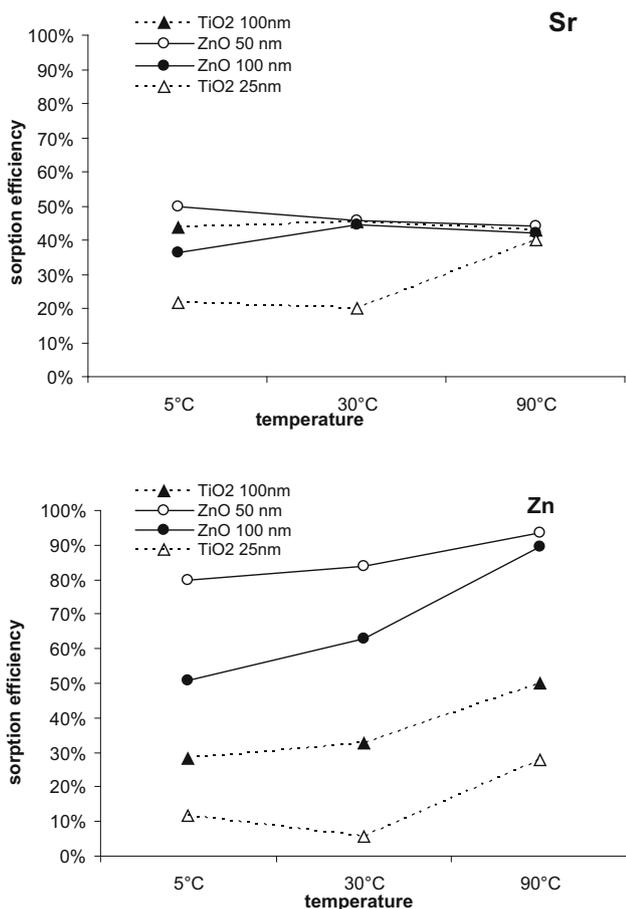


Fig. 2 Influence of temperature on the sorption efficiency of Sr and Zn for TiO₂ 25 nm, TiO₂ 100 nm, ZnO 50 nm, and ZnO 100 nm NPs (SD ≤ 5%)

onto TiO₂ NPs was higher or comparable with the results recorded for pH 8. Similar relation was observed for Zn sorption onto ZnO NPs, while the sorption of Sr under the same conditions was significantly reduced to about 62 and 67% for ZnO 50 nm and ZnO 100 nm, respectively. Observed difference in the behavior of Sr and Zn ions at relatively higher pH can be explained by their chemical character Sr had alkaline—whereas Zn had amphoteric properties. The chemistry of the adsorbent surface is also pH dependent. For the lowest pH, we can expect that some dissolution of NPs occurs. Additionally, the sorbent surface is covered by H⁺ ions. Consequently, the efficiency of sorption is negligible. Increase of pH results in change in the surface properties of NPs; finally, in alkaline pH, there are reactions of hydroxide ions with hydrous metal oxides. Detailed discussion of the possible processes occurring on the metal oxide nanoparticle surface as an effect of pH change can be found in the work of Sheela et al. (2012).

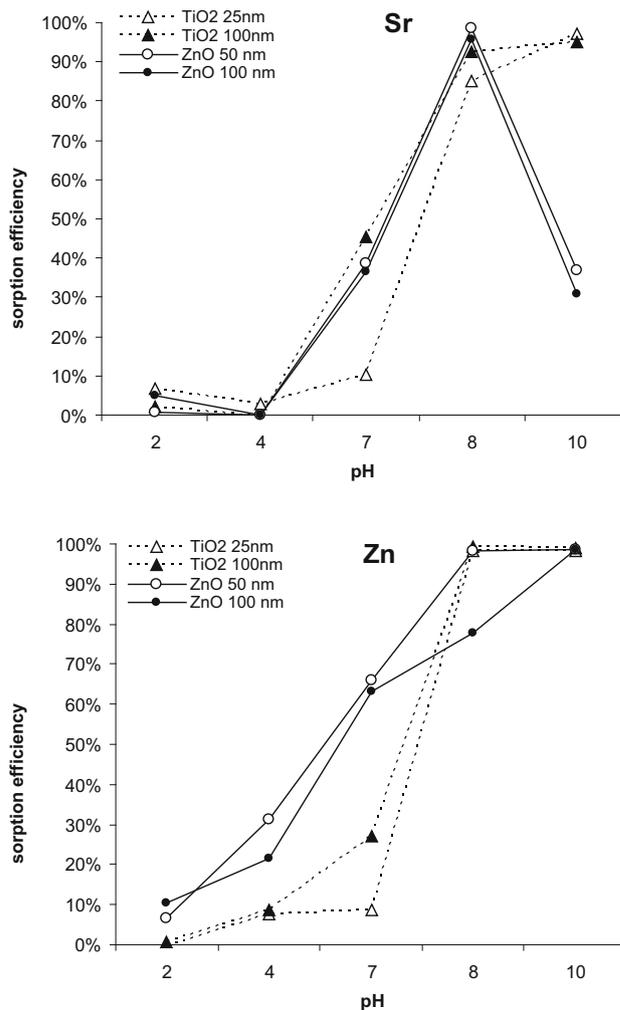


Fig. 3 Influence of pH on the sorption efficiency of Sr and Zn for TiO₂ 25 nm, TiO₂ 100 nm, ZnO 50 nm, and ZnO 100 nm NPs (SD ≤ 5%)

Characteristics of Sorption: Effect of Humic Acids

The above-presented result was not in the context of finding optimum conditions for the sorption of Zn and Sr ions onto metal oxide nanoparticles, but mainly based on the characteristic of sorption under different conditions from the perspective of the processes and interactions, which can take place in the natural aquatic environment affecting the bioaccumulation of the metal ions. To complete the study, the influence of organic matter on the sorption efficiency was examined using humic acids present in natural waters as example compounds (Fig. 4). The results have shown that Sr and Zn sorption onto both investigated ZnO NPs were not affected by the humic acids at applied concentration alike, at short (15 min) as well as long (24 h) interactions. Different observation was found for TiO₂ NPs. Sorption efficiency of Sr on TiO₂ 25 nm increased from about 10 to 40–50% after addition of humic

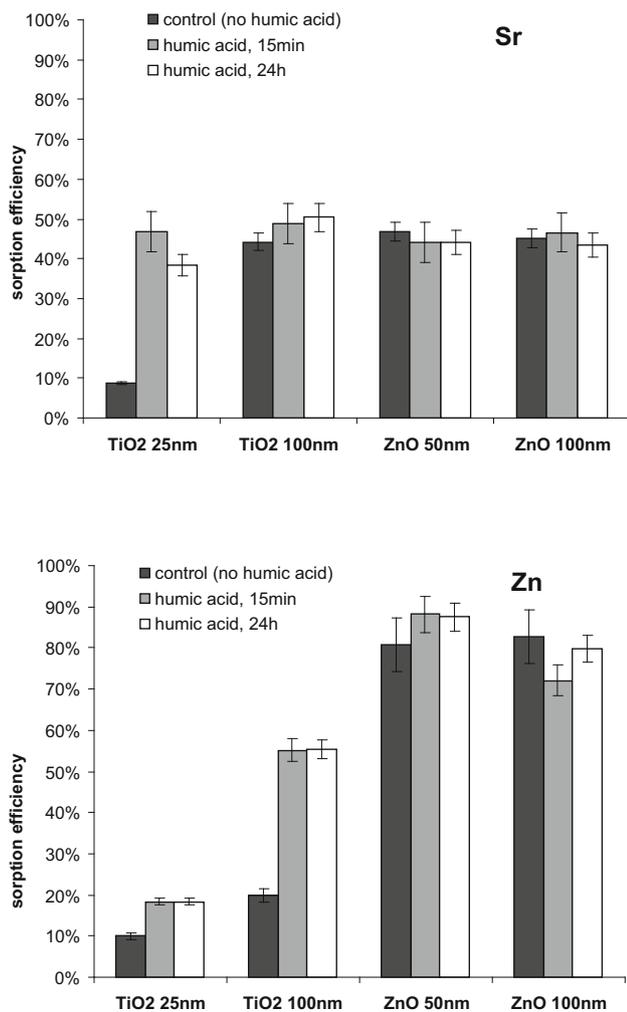


Fig. 4 Influence of humic acids on the sorption efficiency of Sr and Zn for TiO₂ 25 nm, TiO₂ 100 nm, ZnO 50 nm and ZnO 100 nm NPs

acids. For Zn ions, this increase was much lower, but still significant (from 10 to 18%). Definitely, a greater effect was observed for Zn sorption onto TiO₂ 100 nm; addition of humic acids caused the increase of sorption efficiency from 20 to 50%. Sorption of Sr onto NPs of TiO₂ 100 nm was not affected by humic acids. NPs have a strong tendency to agglomerate. The presence of organic matter provides a means of controlling the state of aggregation by forming a charged external interaction with the nanoparticle surface and improving the steric stability of the system (Grillo et al. 2015). After having interacted with humic acids NPs aggregates could be broken down and the hydrodynamic diameter of NPs should be much lower than the aggregate state (Wang et al. 2016a, b). Increased sorption capacity of nanoparticles in the presence of humic acids may be due to: increased available surface area resulting from higher dispersion and disaggregation and increasing electrostatic attraction. The results again confirmed that sorption depends on the chemical composition

of the medium, which affects the chemical properties of both ions and sorbents.

Effect of Nanoparticles on Bioaccumulation of Divalent Ions in Plants

The presented study has undoubtedly shown that NPs can adsorb metal ions present in the water. It also shows that sorption strongly depends on the pH, chemical composition, temperature, or concentrations of NPs, and that under real conditions sorption will be different for different aquatic ecosystems. The question of whether sorption can affect bioavailability of some essential or toxic elements is still open.

The final step of the presented study focused on the examination of the influence of the metal oxide nanoparticles on the bioaccumulation of selected divalent metal ions by aquatic plants: *S. natans* and *E. canadensis*. Results are presented in Fig. 5.

Bioaccumulation of Sr in *E. canadensis* was reduced by NPs of TiO₂ 100 nm, ZnO 50 nm, and ZnO 100 nm at concentrations (1 g L⁻¹) of about 38, 47, and 49%, respectively. Nanoparticles of TiO₂ 25 nm have negligible effect. The statistical analysis of the results of bioaccumulation of Sr in *E. canadensis* in the presence of NPs at 0.1 g L⁻¹ has shown that differences between particular results are statistically insignificant. Thus, bioavailability of Sr under applied conditions was not affected by NPs. The results of Sr accumulation in *S. natans* were similar. Lower concentration (0.1 g L⁻¹) of investigated NPs has not affected the bioaccumulation of Sr; NPs of TiO₂ 25 nm, TiO₂ 100 nm, ZnO 50 nm, and ZnO 100 nm at 1 g L⁻¹ of concentration have reduced the accumulation by 40, 26, 22, and 23%, respectively. Bioavailability of Zn was affected the most by NPs of ZnO. *E. canadensis* accumulated about 45 and 55% less ZnO 50 nm and ZnO 100 nm, respectively, at 0.1 g L⁻¹. Observed reduction in the bioaccumulation of Zn was significantly higher for NPs at 1 g L⁻¹ concentration and amounted to 77 and 68% for ZnO 50 nm and ZnO 100 nm, respectively. Some amount of reduction in the bioavailability of Zn by TiO₂ 25 nm NPs at 0.1 and 1 g L⁻¹ concentrations was observed (22 and 48%, respectively); however, no effect was noticed for TiO₂ 100 nm. Definitely, a different effect on the bioaccumulation of Zn was found for *S. natans*. Addition of ZnO NPs caused an increase of Zn bioaccumulation by 92 and 45% for ZnO 50 nm and ZnO 100 nm, respectively. No effect was observed for both the investigated TiO₂ NPs at 0.1 g L⁻¹ and for all investigated NPs at 1 g L⁻¹.

The presented results show that NPs affect the bioaccumulation of metal ions. Predominant reduction in the metal ion uptake was observed. The effect depends on the chemical composition, size and concentration of NPs, and

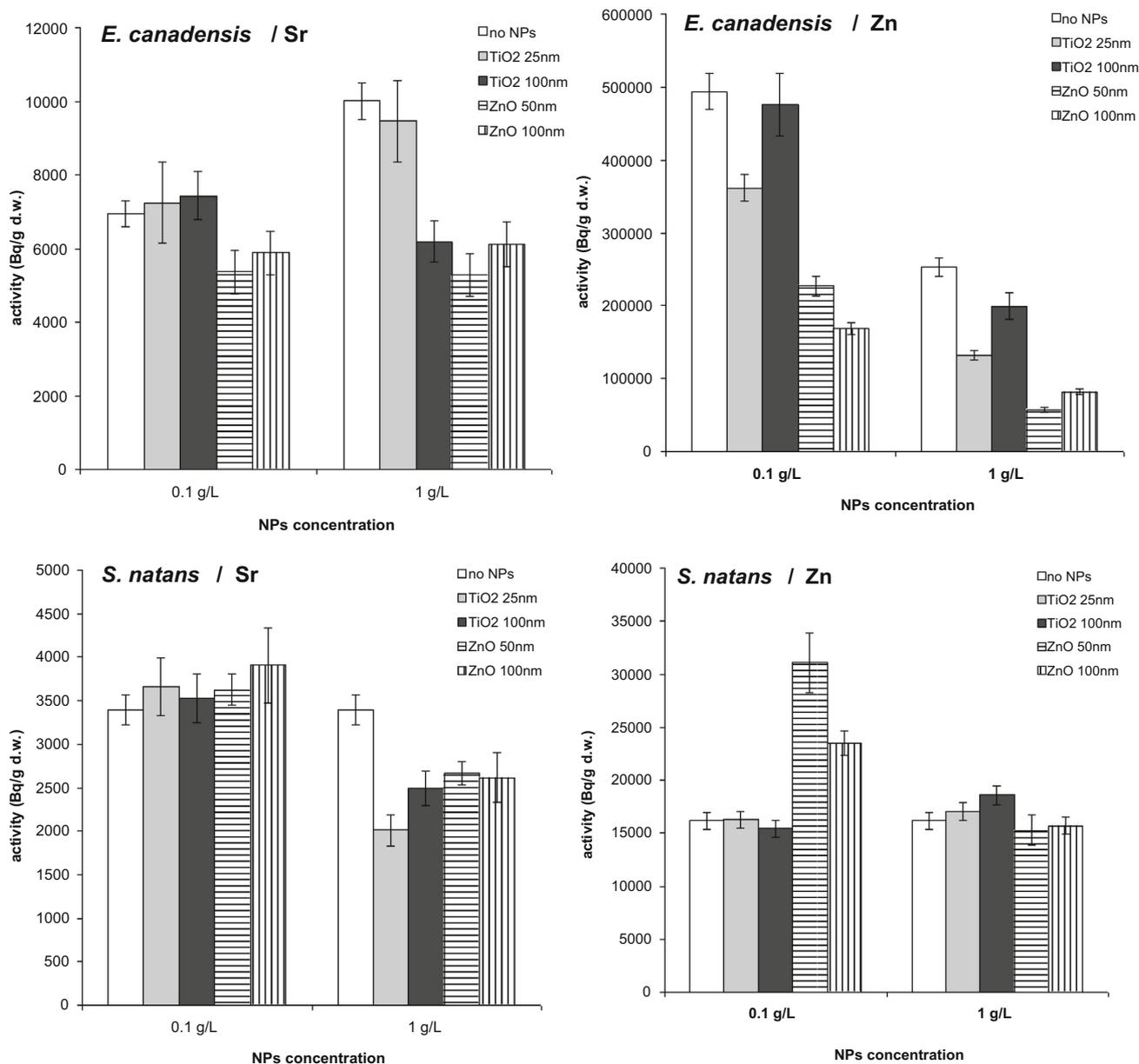


Fig. 5 Bioaccumulation of Sr and Zn by *S. natans* and *E. canadensis* in the presence of TiO₂ 25 nm, TiO₂ 100 nm, ZnO 50 nm, and ZnO 100 nm NPs

the chemical properties of the ions. As the results of metal ion sorption onto NPs have shown, under some conducive conditions, even 90% of ions can be removed from the solution by sorption processes. In consequence, the bioavailability of ions is reduced. Adsorption of NPs onto the root surface, resulting in the inhibition of uptake, can also be the possible mechanism for the reduction of bioaccumulation. However, sorption of ions onto NPs previously adsorbed on the root surface or uptake of NPs together with ions adsorbed cannot be excluded. Another possibility can be the influence of NPs on the transport systems in plants by the formation of new and large-sized

pores and routes, causing an increase of ion uptake. It is also probable that several of the mentioned mechanisms by NPs, which modify the bioaccumulation of metal ions, occur. The next important factor affecting the influence of NPs on the bioaccumulation of ions by plants is the variability in uptake mechanism, which is dependent on physiological, anatomical, and morphological diversities among particular plants species. In our study, we have employed two different plant species: *E. canadensis*, which is a perennial aquatic plant that lives entirely underwater and *S. natans*, which is an annual aquatic floating fern. *S. natans* has the ability to change the pH of the water nearby

the roots. As it can be seen, from the results presented in Fig. 5, the concentration of the accumulated Zn and Sr was much higher in case of *E. canadensis*. These differences between the plant species explain the different observed effects of NPs on bioaccumulation of ions.

Summary

The influence of NPs on the chemical and physical behavior of water components should be considered during the risk assessment of NP contamination of aquatic environment. One of the possible effects caused by NPs is the modification of metal-ion uptake by direct NP–metal ion reaction, including adsorption of ions onto the surface of NPs. Our study showed that NPs of titanium and zinc oxides have significant sorption capacity toward divalent ions of Zn and Sr. Influence of particle size and concentration, sorption temperature, pH, and time on the adsorption efficiency was documented. The examination of the bioaccumulation of Zn(II) and Sr(II) ions by *S. natans* and *E. canadensis* in the presence of metal oxide nanoparticles has led to the conclusion that NPs affect the bioaccumulation of metal ions. The effect depends on the chemical composition, size, and concentration of NPs, and the chemical properties of the ions and plant species. Sorption of ions onto the surface of NPs is one of the possible mechanisms of ion bioavailability reduction. Nevertheless, some other mechanisms are also possible. The results have shown that NPs can modify the bioaccumulation of ions. The effect is positive if ions are toxic, but it is negative if the uptake of essential ions is reduced. Further, more detailed studies must be undertaken to explain the mechanisms responsible for the modification of ion uptake by NPs in aquatic plants.

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