Ionic conductance of H_2Se

Table 1.7.2 Ionic conductances of aqueous solutions

Electrolyte	k or Λ or Λ_0[1] [$\Omega^{-1}cm^{-1}$ or $\Omega^{-1}mol^{-1}cm^2$]	T [K or °C]	c[2] [mol · dm^{-3}]	k or Λ or Λ_0[1] [$\Omega^{-1}cm^{-1}$ or $\Omega^{-1}mol^{-1}cm^2$]	T [K or °C]	c[2] [mol · dm^{-3}]	Ref.
H_2Se	*119.00*	25 °C	0.00408	*92.00*	25 °C	0.00711	[23deH]
	74.40	25 °C	0.01099	*70.60*	25 °C	0.01247	[23deH]
	35.80	25 °C	0.10000	*844*	25 °C		[23deH]

[1]Conductances at infinite dilution Λ_0 are printed in italics without mentioning a concentration, units are Ω^{-1} $mol^{-1}cm^2$. Molar conductances are given in italics with a concentration value, units are Ω^{-1} $mol^{-1}cm^2$. Simple conductivities are given stating the concentration, units are Ω^{-1} cm^{-1}

[2]Concentrations are molar (units: mol · dm^{-3}), molal concentrations are given in italics (units: mol · kg^{-1}), other concentrations as specified

Symbols and Abbreviations

Short form	Full form
κ, Λ	ionic conductivity
T	temperature
Λ_0	ionic conductance at infinite dilution
c	molar concentration

References

[23deH] de Hłasko, M.: J. Chim. Physiq. **20** (1923) 166.

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