

Ionic conductance of H₂GeO₃

Table 1.7.2 Ionic conductances of aqueous solutions

Electrolyte	<i>k</i> or Λ or Λ_0^1 [$\Omega^{-1}\text{cm}^{-1}$ or $\Omega^{-1}\text{mol}^{-1}\text{cm}^2$]	<i>T</i> [K or °C]	<i>c</i> ² [mol · dm ⁻³]	<i>k</i> or Λ or Λ_0^1 [$\Omega^{-1}\text{cm}^{-1}$ or $\Omega^{-1}\text{mol}^{-1}\text{cm}^2$]	<i>T</i> [K or °C]	<i>c</i> ² [mol · dm ⁻³]	Ref.
H ₂ GeO ₃	<i>1.2640</i>	20 °C	0.00880	<i>0.9880</i>	20 °C	0.01465	[32Sch]
	<i>0.0606</i>	25 °C	0.04513	from hydrolysis			[36Lan]
	<i>0.0636</i>	25 °C	0.05413	from hydrolysis			[36Lan]
	<i>0.0578</i>	25 °C	0.04513	from hydrolysis			[36Lan]
	<i>0.0595</i>	25 °C	0.04453	glassy			[36Lan]
	<i>0.101</i>	25 °C	0.05277	heated to 1050 °C			[36Lan]
	<i>0.0970</i>	25 °C	0.04632	from sulfide			[36Lan]

¹Conductances at infinite dilution Λ_0 are printed in italics without mentioning a concentration, units are $\Omega^{-1}\text{mol}^{-1}\text{cm}^2$. Molar conductances are given in italics with a concentration value, units are $\Omega^{-1}\text{mol}^{-1}\text{cm}^2$. Simple conductivities are given stating the concentration, units are $\Omega^{-1}\text{cm}^{-1}$

²Concentrations are molar (units: mol · dm⁻³), molal concentrations are given in italics (units: mol · kg⁻¹), other concentrations as specified

Symbols and Abbreviations

Short form	Full form
κ , Λ	ionic conductivity
<i>T</i>	temperature
Λ_0	ionic conductance at infinite dilution
<i>c</i>	molar concentration

References

[32Sch] Schwarz, R., Huf, E.: Z. Anorg. Allgem. Chem. **203** (1932) 210.

[36Lan] Landolt-Börnstein: Physikalisch-Chemische Tabellen, Erg.-Band IIIc 1936.