

Ionic conductance of copper bimalonate

Table 1.7.2 Ionic conductances of aqueous solutions

Electrolyte	k or Λ or Λ_0^1 [$\Omega^{-1}\text{cm}^{-1}$ or $\Omega^{-1}\text{mol}^{-1}\text{cm}^2$]	T [K or °C]	c^2 [$\text{mol} \cdot \text{dm}^{-3}$]	k or Λ or Λ_0^1 [$\Omega^{-1}\text{cm}^{-1}$ or $\Omega^{-1}\text{mol}^{-1}\text{cm}^2$]	T [K or °C]	c^2 [$\text{mol} \cdot \text{dm}^{-3}$]	Ref.
Cu(C ₃ H ₃ O ₄) ₂	<i>240.0</i>	25 °C	0.0012500	<i>204.8</i>	25 °C	0.0025000	[35Bri]
Copper	<i>165.6</i>	25 °C	0.0050000	<i>131.2</i>	25 °C	0.0100000	[35Bri]
bimalonate	<i>100.4</i>	25 °C	0.0200000	<i>227.6</i>	25 °C		[35Bri]

¹Conductances at infinite dilution Λ_0 are printed in italics without mentioning a concentration, units are $\Omega^{-1} \text{mol}^{-1} \text{cm}^2$. Molar conductances are given in italics with a concentration value, units are $\Omega^{-1} \text{mol}^{-1} \text{cm}^2$. Simple conductivities are given stating the concentration, units are $\Omega^{-1} \text{cm}^{-1}$

²Concentrations are molar (units: $\text{mol} \cdot \text{dm}^{-3}$), molal concentrations are given in italics (units: $\text{mol} \cdot \text{kg}^{-1}$), other concentrations as specified

Symbols and Abbreviations

Short form	Full form
κ, Λ	ionic conductivity
T	temperature
Λ_0	ionic conductance at infinite dilution
c	molar concentration

References

[35Bri] Britton, H.T.S., Jarrett, M.D.: J. Chem. Soc. **1935** (1935) 168.