

Ionic conductance of $\text{CoN}_6\text{Br}_3\text{H}_{18}$

Table 1.7.2 Ionic conductances of aqueous solutions

Electrolyte	k or Λ or Λ_0^1 [$\Omega^{-1}\text{cm}^{-1}$ or $\Omega^{-1}\text{mol}^{-1}\text{cm}^2$]	T [K or °C]	c^2 [$\text{mol} \cdot \text{dm}^{-3}$]	k or Λ or Λ_0^1 [$\Omega^{-1}\text{cm}^{-1}$ or $\Omega^{-1}\text{mol}^{-1}\text{cm}^2$]	T [K or °C]	c^2 [$\text{mol} \cdot \text{dm}^{-3}$]	Ref.
$\text{CoN}_6\text{Br}_3\text{H}_{18}$	<i>90.6</i>	0 °C	0.00050	<i>88.4</i>	0 °C	0.00100	[29Int]
	<i>85.7</i>	0 °C	0.00200	<i>80.2</i>	0 °C	0.00500	[29Int]
	<i>75.0</i>	0 °C	0.01000	<i>69.9</i>	0 °C	0.02000	[29Int]
	<i>63.9</i>	0 °C	0.05000	<i>60.0</i>	0 °C	0.07000	[29Int]
	<i>56.7</i>	0 °C	0.10000				[29Int]

¹Conductances at infinite dilution Λ_0 are printed in italics without mentioning a concentration, units are $\Omega^{-1} \text{mol}^{-1} \text{cm}^2$. Molar conductances are given in italics with a concentration value, units are $\Omega^{-1} \text{mol}^{-1} \text{cm}^2$. Simple conductivities are given stating the concentration, units are $\Omega^{-1} \text{cm}^{-1}$

²Concentrations are molar (units: $\text{mol} \cdot \text{dm}^{-3}$), molal concentrations are given in italics (units: $\text{mol} \cdot \text{kg}^{-1}$), other concentrations as specified

Symbols and Abbreviations

Short form	Full form
κ, Λ	ionic conductivity
T	temperature
Λ_0	ionic conductance at infinite dilution
c	molar concentration

References

[29Int] International critical tables of numerical data, physics, chemistry and technology Vol. 6 MacGraw-Hill, New York 1929.