

Ionic conductance of Cd(NO₃)₂+LiNO₃

Table 1.7.2 Ionic conductances of aqueous solutions

Electrolyte	<i>k</i> or Λ or Λ_0^1 [$\Omega^{-1}\text{cm}^{-1}$ or $\Omega^{-1}\text{mol}^{-1}\text{cm}^2$]	<i>T</i> [K or °C]	<i>c</i> ² [mol · dm ⁻³]	<i>k</i> or Λ or Λ_0^1 [$\Omega^{-1}\text{cm}^{-1}$ or $\Omega^{-1}\text{mol}^{-1}\text{cm}^2$]	<i>T</i> [K or °C]	<i>c</i> ² [mol · dm ⁻³]	Ref.
Cd(NO ₃) ₂ + LiNO ₃	0.14494	25 °C	0.0 + 3.0	0.13504	25 °C	0.375 + 2.25	[38Rys]
	0.12613	25 °C	0.75 + 1.5				[38Rys]
	0.11741	25 °C	1.125 + 0.75	0.10881	25 °C	1.5 + 0.0	[38Rys]
	0.16910	25 °C	0.0 + 5.0	0.14894	25 °C	0.626 + 3.75	[38Rys]
	0.13345	25 °C	1.25 + 2.5	0.11797	25 °C	1.875 + 1.25	[38Rys]
	0.10432	25 °C	2.0 + 0.0				[38Rys]

¹Conductances at infinite dilution Λ_0 are printed in italics without mentioning a concentration, units are $\Omega^{-1}\text{mol}^{-1}\text{cm}^2$. Molar conductances are given in italics with a concentration value, units are $\Omega^{-1}\text{mol}^{-1}\text{cm}^2$. Simple conductivities are given stating the concentration, units are $\Omega^{-1}\text{cm}^{-1}$

²Concentrations are molar (units: mol · dm⁻³), molal concentrations are given in italics (units: mol · kg⁻¹), other concentrations as specified

Symbols and Abbreviations

Short form	Full form
κ , Λ	ionic conductivity
<i>T</i>	temperature
Λ_0	ionic conductance at infinite dilution
<i>c</i>	molar concentration

References

[38Rys] Rysselberge, P. van, Lee, A.: J. Am. Chem. Soc. **60** (1938) 2776.