Ionic conductance of acetonitrile/isobutyronitrile-based electrolyte solutions

Table 1.7.9 Ionic conductances of mixed solvent-based electrolyte solutions

Solvent[1]	Salt	k or Λ or Λ_0[2] [$\Omega^{-1}cm^{-1}$ or $\Omega^{-1}mol^{-1}cm^2$]	T [K]	c[3] [mol · dm^{-3}]	Ref.
Acetonitrile/isobutyronitrile 20.4/79.6 % wt.	Bu_4NBPh_4	*81.33*	25 °C	14.682 10^{-4}	[60Bro]
		79.53	25 °C	20.670 10^{-4}	[60Bro]
		78.86	25 °C	23.532 10^{-4}	[60Bro]
		77.28	25 °C	30.889 10^{-4}	[60Bro]
		75.88	25 °C	38.610 10^{-4}	[60Bro]
		92.77	25 °C		[60Bro]

[1]Composition in units as stated
[2]Conductances at infinite dilution Λ_0 are printed in italics without mentioning a concentration, units are Ω^{-1} $mol^{-1}cm^2$. Molar conductances are given in italics with a concentration value, units are Ω^{-1} $mol^{-1}cm^2$. Simple conductivities are given stating the concentration, units are Ω^{-1} cm^{-1}
[3]Concentrations are molar (units: mol · dm^{-3}), molal concentrations are given in italics (units: mol · kg^{-1}), other concentrations as specified

Symbols and Abbreviations

Short form	Full form
κ, Λ	ionic conductivity
T	temperature
Λ_0	ionic conductance at infinite dilution
c	molar concentration

References

[60Bro] Brown, A.M., Fuoss, R.M.: J. Phys. Chem. **64** (1960) 1341.

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