

Ionic conductances of EC:DME nonaqueous electrolyte solution of LiN(CF₃SO₂)₂ salt

Table 1.7.6 Ionic conductances of nonaqueous electrolyte solutions

Solvent	Salt	k^1 [$\Omega^{-1}\text{cm}^{-1}$]	T [K]	c^2 [$\text{mol} \cdot \text{dm}^{-3}$]	k^1 [$\Omega^{-1}\text{cm}^{-1}$]	T [K]	c^2 [$\text{mol} \cdot \text{dm}^{-3}$]	Ref.
EC:DME 50:50 %vol.	LiN(CF ₃ SO ₂) ₂	<i>7.87</i>	0 °C	1	<i>12.08</i>	20 °C	1	[11Lin3]
		<i>16.58</i>	40 °C	1	<i>21.25</i>	60 °C	1	[11Lin3]
		<i>25.97</i>	80 °C	1				[11Lin3]

¹Specific conductivities k with units as stated, molar conductivities Λ are printed in italics, molar conductivities at infinite dilutions Λ_0 units are printed in italics without stating a concentration, units $\Omega^{-1}\text{cm}^2\text{mol}^{-1}$

²Concentrations are molar (units: $\text{mol} \cdot \text{dm}^{-3}$), molal concentrations are given in italics (units: $\text{mol} \cdot \text{kg}^{-1}$), other concentrations as specified

Symbols and Abbreviations

Short form	Full form
T	temperature
k	ionic conductance
c	molar concentration

References

[11Lin3] Linden's Handbook of Batteries, Reddy, T.B., Ed., McGraw Hill, New York, 2011.