

## Single ion conductivities of $[\text{Cr}(\text{NH}_3)_5\text{OH}][\text{OH}]_2\text{HCl}$

**Table 1.7.5** Single ion conductivities of ions in aqueous electrolyte solutions as a function of their concentration, entries arranged alphabetically

Electrolyte	$\lambda^+$ [ $\Omega^{-1}\text{m}^2\text{mol}^{-1}$ ]	$\lambda^-$ [ $\Omega^{-1}\text{m}^2\text{mol}^{-1}$ ]	$c^1$ [ $\text{mol} \cdot \text{dm}^{-3}$ ]	$T$ [K]	Ref.
$[\text{Cr}(\text{NH}_3)_5\text{OH}][\text{OH}]_2\text{HCl}$	$78 \cdot 10^{-4}$		0.00135	0 °C	[31Lan]
	$342.5 \cdot 10^{-4}$		0.005	25 °C	[43Sto]
	$399.9 \cdot 10^{-4}$	$76.3 \cdot 10^{-4}$	0.01	25 °C	[43Sto]
	$336.5 \cdot 10^{-4}$	$65.8 \cdot 10^{-4}$	0.02	25 °C	[43Sto]
	$330.9 \cdot 10^{-4}$	$63.1 \cdot 10^{-4}$	0.05	25 °C	[43Sto]
	$325.2 \cdot 10^{-4}$	$60.3 \cdot 10^{-4}$	0.1	25 °C	[43Sto]
	$317.9 \cdot 10^{-4}$	$56.9 \cdot 10^{-4}$	0.2	25 °C	[43Sto]

<sup>1</sup>Concentrations are molar (units:  $\text{mol} \cdot \text{dm}^{-3}$ ), molal concentrations are given in italics (units:  $\text{mol} \cdot \text{kg}^{-1}$ ), other concentrations as specified

### Symbols and Abbreviations

Short form	Full form
$T$	temperature
$\lambda$	ionic conductance
$c$	molar concentration

### References

- [31Lan] Landolt-Börnstein: Physikalisch-Chemische Tabellen, Erg.-Band IIb 1931.  
 [43Sto] Stonehill, H.J.: J. Chem. Soc. **1943** (1943) 647.