Ionic conductance of sodium mellithat

Table 1.7.2 Ionic conductances of aqueous solutions

Electrolyte	k or Λ or Λ_0[1] [$\Omega^{-1}cm^{-1}$ or $\Omega^{-1}mol^{-1}cm^2$]	T [K or °C]	c[2] [mol · dm^{-3}]	k or Λ or Λ_0[1] [$\Omega^{-1}cm^{-1}$ or $\Omega^{-1}mol^{-1}cm^2$]	T [K or °C]	c[2] [mol · dm^{-3}]	Ref.
$Na_4C_{10}H_2O_8$ Sodium mellithat	*480.0*	25 °C	0.0010000	*464.4*	25 °C	0.0020000	[29Int]
	433.2	25 °C	0.0050000	*404.4*	25 °C	0.0100000	[29Int]
	372.8	25 °C	0.0200000	*329.6*	25 °C	0.0500000	[29Int]
	314.0	25 °C	0.0700000				[29Int]

[1]Conductances at infinite dilution Λ_0 are printed in italics without mentioning a concentration, units are Ω^{-1} $mol^{-1}cm^2$. Molar conductances are given in italics with a concentration value, units are Ω^{-1} $mol^{-1}cm^2$. Simple conductivities are given stating the concentration, units are Ω^{-1} cm^{-1}

[2]Concentrations are molar (units: mol · dm^{-3}), molal concentrations are given in italics (units: mol · kg^{-1}), other concentrations as specified

Symbols and Abbreviations

Short form	Full form
κ, Λ	ionic conductivity
T	temperature
Λ_0	ionic conductance at infinite dilution
c	molar concentration

References

[29Int] International critical tables of numerical data, physics, chemistry and technology Vol. 6 MacGraw-Hill, New York 1929.

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